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the ph scale is used to rank solutions in terms of acidity or basicity alkalinity since the scale is based on ph values it is logarithmic meaning that a change of 1 ph unit corresponds to a ten fold change in h start superscript plus end superscript ion concentration the ph scale is often said to range from 0 to 14 and most solutions do fall within this range although it s possible to get a ph below 0 or above 14 ph quantitative measure of the acidity or basicity of aqueous or other liquid solutions the term widely used in chemistry biology and agronomy translates the values of the concentration of the hydrogen ion which ordinarily ranges between about 1 and 10⁻¹⁴ gram equivalents per litre into numbers between 0 and 14 thus ph may be defined as a measure of free acidity more precisely ph is defined as the negative log of the hydrogen ion concentration the range of ph extends from zero to 14 a ph value of 7 is neutral because pure water has a ph value of exactly 7 values lower than 7 are acidic values greater than 7 are basic or alkaline ph log h a water source s ph value is a function of its acidity or alkalinity the ph level is a function of the hydrogen atom activity as the hydrogen activity is a reasonable indicator of the water s acidity or alkalinity as seen below the ph scale ranges from 0 to 14 when 7 0 is neutral may 6 2019 ph is a measure of hydrogen ion concentration a measure of the acidity or alkalinity of a solution the ph scale usually ranges from 0 to 14 aqueous solutions at 25 c with a ph less than 7 are acidic while those with a ph greater than 7 are basic or alkaline in chemistry ph p i? ? e? t? historically denoting potential of hydrogen or power of hydrogen is a scale used to specify the acidity or basicity of an aqueous solution acidic solutions solutions with higher concentrations of h ions are measured to have lower ph values than basic or alkaline solutions noun ?p? ??ch a measure of acidity and alkalinity of a solution that is a

number on a scale on which a value of 7 represents neutrality and lower numbers indicate increasing acidity and higher numbers increasing alkalinity and on which each unit of change represents a tenfold change in acidity or alkalinity and that is the negative logarithm of the effective hydrogen ion concentration or hydrogen ion activity in gram equivalents per liter of the solution may 6 2019 ph is a logarithmic measure of the hydrogen ion concentration of an aqueous solution $\text{pH} = -\log h$ where \log is the base 10 logarithm and h is the hydrogen ion concentration in moles per liter pH describes how acidic or basic an aqueous solution is where a pH below 7 is acidic and a pH greater than 7 is basic pH of mar 2 2019 pH is a measure of how acidic basic water is the range goes from 0 to 14 with 7 being neutral pH s of less than 7 indicate acidity whereas a pH of greater than 7 indicates a base the pH of water is a very important measurement concerning water quality water science school home water properties topics water quality topics pH and water mar 24 2022 $\text{pH} = -\log_{10} h$ the term is used to indicate basicity or acidity of a solution on a scale of 0 to 14 with pH 7 being neutral as the concentration of h ions in solution increases acidity increases and pH gets lower below 7 see figure 1 when pH is above 7 the solution is basic figure 1 example of an acidic stream due to mine drainage

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